

# Stoichiometry Chapter 12 Test A Answers

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

12

## STOICHBOMETRY

### Chapter Test B

#### A. Matching

Match each item in Column A with the correct description in Column B. Write the letter of the correct item on the line.

Column A	Column B
_____ 1. the substance that determines the amount of product that can be formed in a reaction	a. percent yield
_____ 2. the amount of product that forms when a reaction is carried out in the laboratory	b. limiting reagent
_____ 3. the calculation of quantities in chemical equations	c. theoretical yield
_____ 4. the ratio of the actual yield to the theoretical yield expressed as a percent	d. stoichiometry
_____ 5. the substance that is present in enough quantity to react with a limiting reagent	e. excess yield
_____ 6. the maximum amount of product that could be formed from given amounts of reactants	f. excess reagent

#### B. Multiple Choice

Choose the best answer and write its letter on the line.

\_\_\_\_\_ 7. In a chemical reaction

a. atoms are conserved

b. atoms are conserved

c. matter is conserved

d. both mass and volume are conserved

\_\_\_\_\_ 8. Which of the following is a correct designation of the balanced equation?

2H<sub>2</sub>(g) + 2H<sub>2</sub>O<sub>2</sub>(l) → 2H<sub>2</sub>O(l) + O<sub>2</sub>(g)

a. 2 grams H<sub>2</sub> + 2 moles of H<sub>2</sub>O<sub>2</sub> → 2 molecules H<sub>2</sub>O + 2 moles of O<sub>2</sub>

b. 2 grams H<sub>2</sub> + 2 grams H<sub>2</sub>O<sub>2</sub> → 2 grams H<sub>2</sub>O + 2 grams O<sub>2</sub>

c. 2 moles H<sub>2</sub> + 2 moles H<sub>2</sub>O<sub>2</sub> → 2 moles H<sub>2</sub>O + 2 moles O<sub>2</sub>

d. both c and e

e. both a and c

\_\_\_\_\_ 9. If 5.0 moles of HCl are consumed in the reaction below, how many moles of BaCl<sub>2</sub> are produced?

2HCl + Ba(OH)<sub>2</sub> → BaCl<sub>2</sub> + 2H<sub>2</sub>O

a. 0.50 mol

b. 1.0 mol

c. 2.0 mol

d. 4.0 mol

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